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G. Connelly Royal Society of Chemistry (UK) Honesty everyday exam I'll give two exams, one in chemistry and one in honesty. I hope you will pass them both, but if you miss one, let it be chemical, There are many good people in this world today that they cannot spend an exam in chemistry, but there are no good people in the world that cannot spend an examination in the brief. (Variation of a quote from Madison Sarratt (1891-1978), Dean, Dean, University) Prerequisite Knowledge: a non-exhaustive list of incoming expectations for the inorganic chemistry lesson is listed below. The courses of prerequisites are the chemical general I / II General I / II, quantitative analytical chemistry and organic chemistry I / II: Atomic structure (eg. Atomic numbers, mass numbers, isotopes, subatomic particles), dimensional analysis, the periodic table , denomination, ions and ionic compounds, Stotcheometria, of formula weights, masses and moles of interconnection, Termochimica, enthalpy, Behavior of the waves of matter, quantum numbers, quantum mechanics and atomic orbitals, orbital representations (P, D, f), electronic thrust and exclusion principle Electron and Pauli, electron configurations, effective nuclear charge, periodic and periodic trends (size of atoms and ions, affinity of electrons, electronegativity, ionization energy), symbols and Lewis structures, bond and energy ionic, bond and energy ionic, covalent bonding, molecular geometry, VSEPR, bonding theories (the bond Theory Hybridization), covalent bonding and over- orbital ppositione, molecular orbitals for diatomic molecules, the gas laws, phase changes, phase diagrams, solid structures, cell units, close packing of spheres, bonding in solids (molecular network, covalent, ionic, metallic), properties in colligioso , chemical kinetics (EA, read rate), acidity (pH, PKA), titles, balance (based acidatica, solubility, common ion effect), Le Chatelier's principle, classical thermodynamics and the laws of thermodynamics, entropy, Free Gibbs energy , oxidation states, the oxidation reduction reactions of the balance (redox), equation nuclear, nuclear chemistry (eg radioactivity, decay), the crystal field theory, configuration of electrons in complexes, nucleophilic and electrophiles, SN1 and SN2 reactions , spectroscopy (visible, IR, UV, NMR) review of general chemistry I and general chemistry II chemreview.net chemistry review review Activity Geek.net links and chemical resources extra Literature related to inorganic chemistry, inorganic chemistry, 4th ed., Miessler Gary, Donald A. Tarr; Prentice Hall, Inorganic Chemistry in 2010, the 5th Ed., Duaward F. Shriver, Peter W. Atkins, Tina Overton, Jonathan Rourke; W. H. Freeman, 2009 Inorganic Chemistry, 2nd Ed., Duward F. Shriver, Peter W. Atkins, and C. H. Langford; WH Freeman and Co., New York, 1994 Concepts and inorganic chemistry 3rd models, Bodie Eugene Douglas, Hamilton Darl McDaniel, John J. Alexander, Wiley, 1994 Inorganic Chemistry: Principles of structure and reactivity James E. Huheey, Ellen A. Keiter, Richard L. Keiter; Pearson Education, 2000 Inorganic Chemistry Vol 1, 3rd Ed., Duaward F. Shriver, Peter William Atkins; W. H. Freeman and Co., 1999 1999

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